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A solution is a homogeneous mixture of two or more components in which the particle size is smaller than 1 nm.

Common examples of solutions are the sugar in water and salt in water solutions, soda water, etc. In a solution, all the components appear as a single phase.

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There is particle homogeneity i.e. particles are evenly distributed.

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A solution is a homogeneous mixture of two or more components. The dissolving agent is the solvent. The substance that is dissolved is the solute. The components of a solution are atoms, ions, or molecules,

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making them  $10^{-9}$  m or smaller in diameter.

Solutions, Suspensions, Colloids, and  
Dispersions

2 millimoles (mmol) =  $2 \times 10^{-3}$  mol

(0.002 mol) centi: c:  $10^{-2}$ : 7 centimeters

(cm) =  $7 \times 10^{-2}$  m (0.07 m) deci: d:  $10^{-1}$ :

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1 deciliter (dL) =  $1 \times 10^{-1}$  L (0.1 L) kilo:  
k:  $10^3$ : 1 kilometer (km) =  $1 \times 10^3$  m  
(1000 m) mega: M:  $10^6$ : 3 megahertz  
(MHz) =  $3 \times 10^6$  Hz (3,000,000 Hz)  
giga: G:  $10^9$ : 8 gigayears (Gyr) =  $8 \times 10^9$   
9 yr (8,000,000,000 Gyr) tera: T:  $10^{12}$

Measurements | Chemistry

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Solution, in chemistry, a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility. The term solution is

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Dimension 2 commonly applied to the liquid state of matter, but solutions of gases and solids are possible.

solution | Definition & Examples |  
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Solved: A rectangular block has

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